

Name: _____

Period: _____

HW3:9 Test Review
Mr. Murray, IPC
cstephenmurray.com

Assigned: Tues., 10/31 and Wed., 11/1
Due: Thurs., 11/2 and Fri., 11/3

- For Oxygen give the following:
 - Oxidation # _____
 - # of protons _____
 - Valence Electrons _____
 - Atomic Number: _____
 - Draw Oxygen's Lewis Dot Diagram:
 - What is the total charge for:
 - Be_2^{2+} _____
 - $\text{Al}^{3+} \text{O}_2^{2-}$ _____
 - $\text{Li}_2^+ \text{N}^{3-}$ _____
 - A dissolved compound that allows electricity to pass:
 - Using electron arrows, combine Lithium and Oxygen.
 - Write the balanced ionic compound formula for:
 - Mg^{2+} and S^{2-} _____
 - Al^{3+} and Se^{2-} _____
 - Ca and Br _____
 - Al and F _____
 - Na^{1+} and $(\text{CrO}_4)^{2-}$ _____
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6. What elements don't combine (give the family name)?
7. What elements have a varied # of valence electrons?
8. Give the type of compound and name for the following:
P₂Cl₃ _____
Mg(NO₃)₂ _____
CaBr₂ _____
9. Draw the Lewis Dot Diagram to show how Sulfur and Chlorine will combine:

10. How many valence electrons do each of these have?



11. Using shorthand (line bonds), combine Nitrogen and Chlorine in the correct ratio.

Be sure to study your bellwork for the test, too. Study all of the bellwork since the last test. Read your worksheet notes. Find time to use the website

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